



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE NAME

CENTRE NUMBER

CANDIDATE NUMBER

* 5 4 8 3 5 4 7 5 9 5 *

CO-ORDINATED SCIENCES

0654/31

Paper 3 (Extended)

October/November 2012

2 hours

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 32.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use	
1	
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12	
Total	

This document consists of **29** printed pages and **3** blank pages.



- 1 (a) Complete Table 1.1 by choosing one of the words from the list to match statement.

ammeter ampere circuit coulomb electron
ohm relay volt voltmeter watt

Table 1.1

statement	word
a complete loop of conductors	
the unit of electrical charge	
an instrument that measures potential difference	
a device used in switching on circuits	

[2]

- (b) Fig. 1.1 shows two circuits **A** and **B**. All the lamps and both cells are the same.

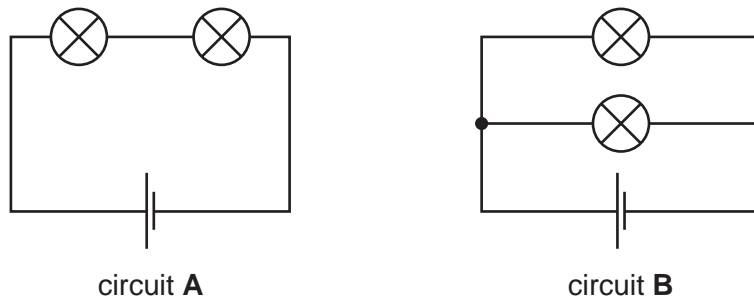


Fig. 1.1

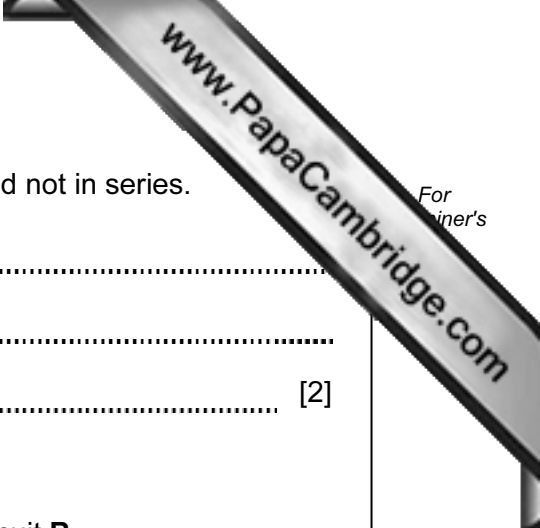
- (i) One lamp is unscrewed from circuit **A**.

State what happens to the other lamp.

Explain your answer.

.....

[1]



(ii) Explain why lights in a house are connected in parallel and not in series.

.....
.....
..... [2]

(iii) The resistance of each lamp is 1.2Ω .

Calculate the combined resistance of the two lamps in circuit **B**.

State the formula that you use and show your working.

formula used

working

..... [3]

2 (a) Fig. 2.1 shows part of the carbon cycle.

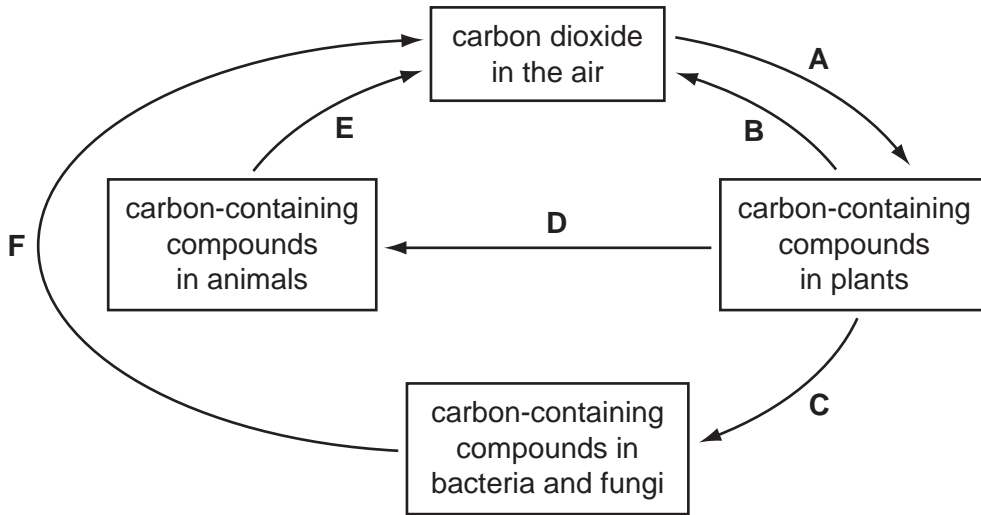


Fig. 2.1

(i) State the letter or letters, **A, B, C, D, E** or **F**, that represent
 photosynthesis,

respiration.

[2]

(ii) Name **one** carbon-containing compound in plants.

..... [1]

(iii) State the approximate percentage of carbon dioxide in the air.

..... [1]

(b) Earthworms play an important part in the carbon cycle. They are decomposers.

Describe the role of decomposers in the carbon cycle.

.....

.....

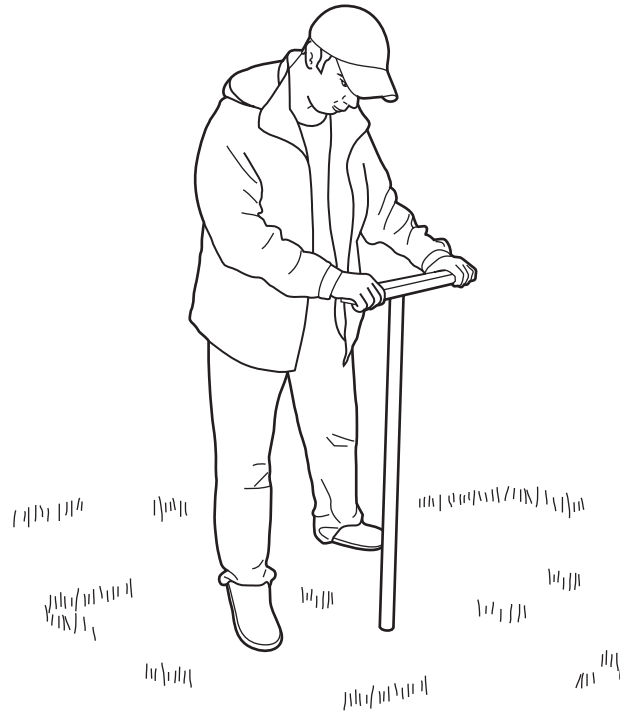
.....

..... [2]

(c) In Florida, USA, some people collect earthworms by vibrating the soil.

A wooden post is pushed into the ground, and then a heavy object is pulled across the top of the post to make it vibrate. The vibrations travel through the soil.

Earthworms respond to the vibrations by crawling out of their burrows onto the soil surface, where they can be caught.



A student investigated the effect of different frequencies of vibrations on the numbers of earthworms that emerged from the soil. Fig. 2.2 shows his results.

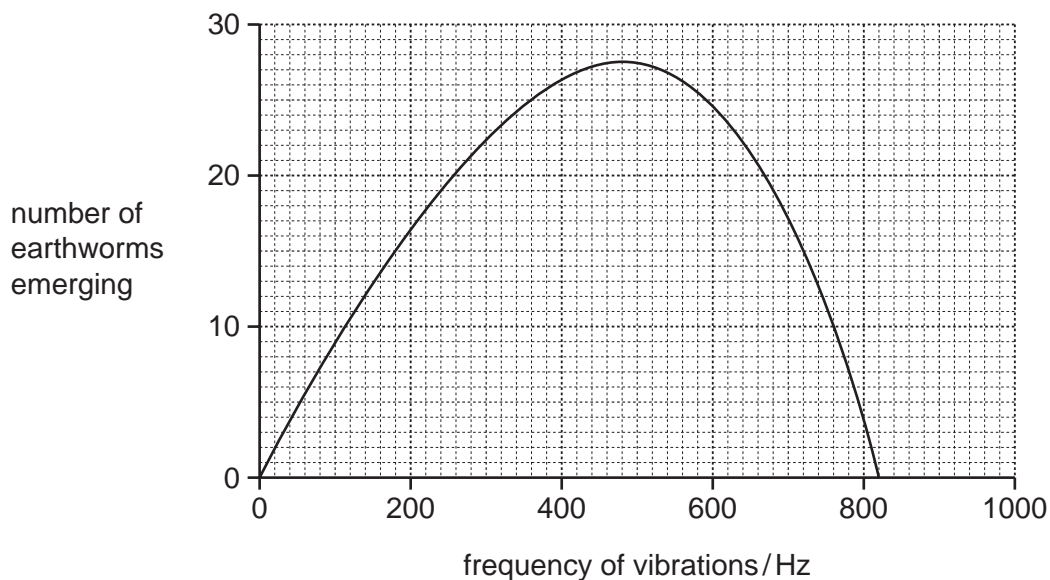
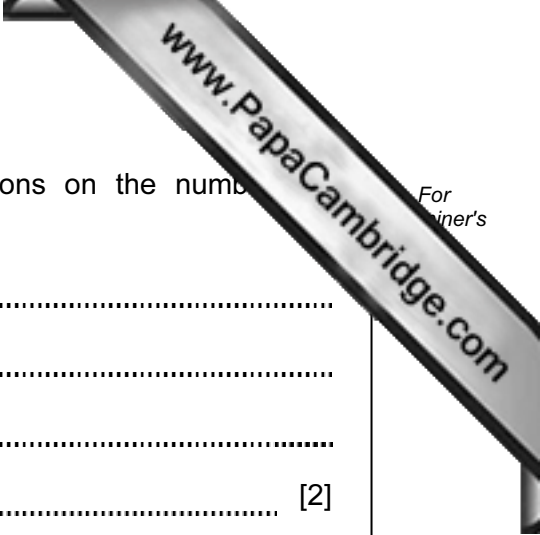


Fig. 2.2



- (i) Describe the effect of different frequencies of vibrations on the number of earthworms emerging.

.....

.....

.....

..... [2]

- (ii) Moles are predators that live underground and eat earthworms. When moles burrow through the ground, they produce vibrations of around 500 Hz.

The response of earthworms to vibrations is controlled by their genes.

Suggest how natural selection may have caused the response of earthworms to vibrations to evolve.

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..... [4]

- 3 (a) Fig. 3.1 shows how a digital pH meter is used to measure the pH of some liquids.

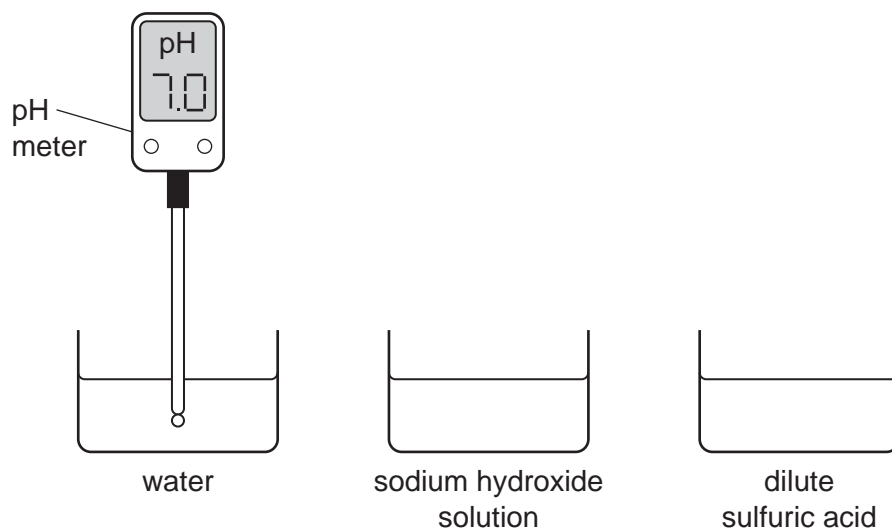


Fig. 3.1

- (i) Complete Table 3.1 by suggesting suitable pH values for the different liquids.

Table 3.1

liquid	pH
water	7.0
sodium hydroxide solution	
dilute sulfuric acid	

[1]

- (ii) Suggest **one** advantage of using a digital pH meter rather than a piece of litmus paper to assess the acidity of an aqueous solution.

.....

..... [1]

- (iii) Dilute acids are aqueous solutions that contain dissolved ions.

Table 3.2 shows the names of the ions in two common acids.

Table 3.2

name of dilute acid	names of dissolved ions
hydrochloric acid	hydrogen ions and chloride ions
sulfuric acid	hydrogen ions and sulfate ions

A student is given an unlabelled beaker which is known to contain either dilute hydrochloric acid or dilute sulfuric acid.

Describe a chemical test that a student could use to find out which acid the beaker contains.

.....

 [2]

- (b) When a reactive metal is added to a dilute acid, the metal reacts and dissolves and hydrogen gas is given off.
- (i) When magnesium reacts with dilute hydrochloric acid, magnesium **atoms** are oxidised by hydrogen **ions**.

The balanced ionic equation for this redox reaction is shown below.

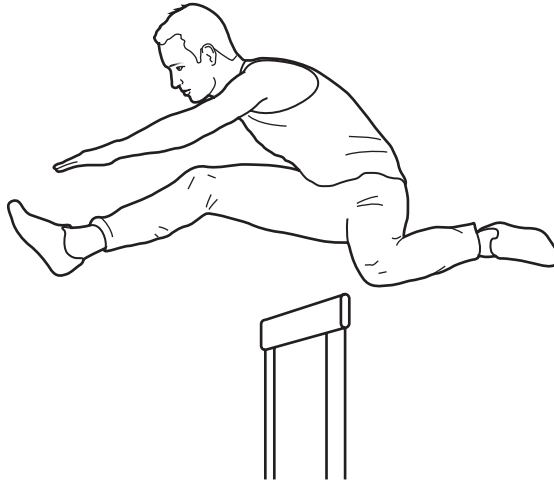


Explain, in terms of the transfer of electrons, why this reaction is described as redox.

.....

 [2]

4 (a) An athlete of mass 60 kg jumps 1.3 metres vertically.



Calculate the work done by the athlete to achieve this height.

State the formula that you use and show your working. The gravitational field strength of the Earth is 10 N/kg.

formula used

working

..... [3]

(b) Using your answer to part (a), state the gain in potential energy of the athlete when he jumps 1.3 metres.

..... [1]

(c) The work done in jumping vertically was completed in 0.5 s.

Calculate the power developed.

State the formula that you use and show your working.

formula used

working

..... [2]

- 5 Fig. 5.1 shows apparatus that can be used to measure the rate of respiration of germinating seeds.

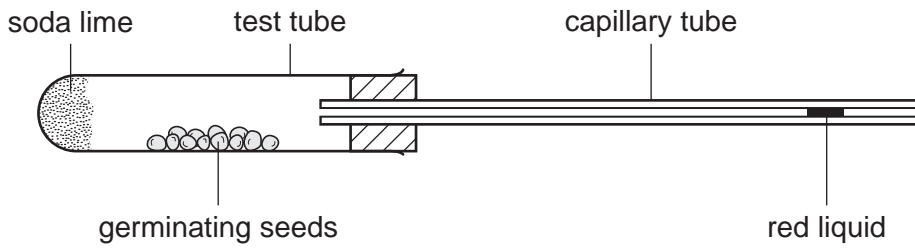


Fig. 5.1

The soda lime absorbs carbon dioxide from the air inside the apparatus.

- (a) As the seeds respire, they use oxygen. This reduces the volume of gas inside the apparatus. The faster they respire, the faster the red liquid moves towards the left.

- (i) Write the balanced equation for aerobic respiration.

..... [2]

- (ii) Use the equation to explain why the liquid would **not** move if there was **no** soda lime in the apparatus.

.....

 [2]

- (b) An experiment was carried out to investigate the effect of temperature on the respiration of the germinating seeds.

Four sets of the apparatus shown in Fig. 5.1 were set up and labelled **A**, **B**, **C** and **D**. Each set of apparatus contained either germinating or dead seeds.

The distance moved by the red liquid in five minutes was measured for each set.

The results are shown in Table 5.1.

Table 5.1

set	contents	temperature / °C	distance moved by red liquid in 5 minutes / mm
A	germinating seeds	0	3
B	germinating seeds	10	6
C	germinating seeds	20	12
D	dead seeds	20	0

- (i) Explain why it was important to include set **D** in the experiment.

.....
..... [1]

- (ii) Suggest why the liquid may have moved very slightly in set **D**.

.....
..... [1]

- (iii) With reference to Table 5.1, describe the effect of temperature on the rate of respiration of germinating seeds.

.....
.....
..... [2]

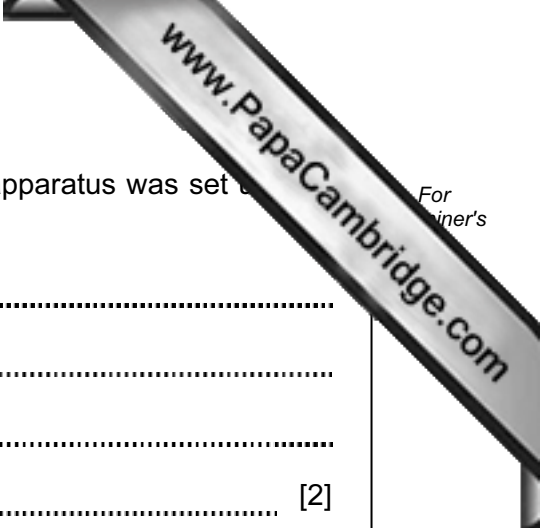
- (iv) Predict and explain the results you would expect if the apparatus was set up to germinating seeds at a temperature of 60 °C.

predicted results

explanation

.....

..... [2]



- 6 Some types of firework are made by filling a cardboard tube with firework mixture. The mixture is made from several solid substances which have been powdered and mixed together.

Fig. 6.1 shows a typical firework.

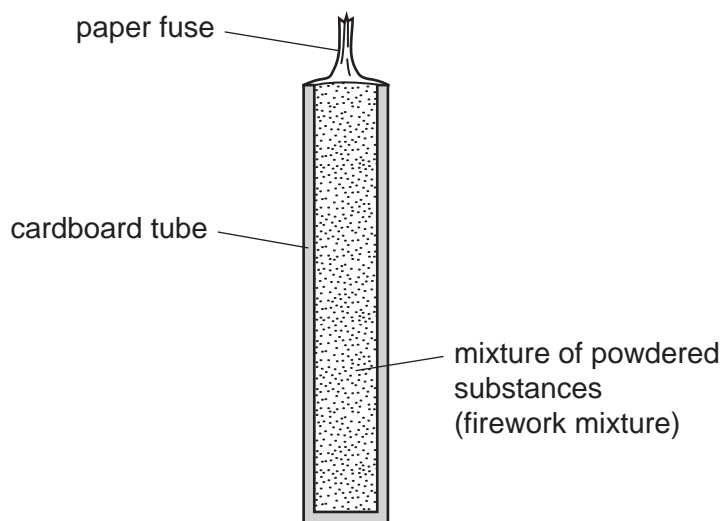


Fig. 6.1

When the paper fuse is lit, exothermic chemical reactions occur inside the firework.

- (a) Explain, in terms of rate of reaction, why firework mixture is a powder.

.....

 [2]

- (b) Some firework mixtures contain aluminium which is oxidised to produce aluminium oxide.

When aluminium is oxidised, aluminium atoms are converted into aluminium ions.

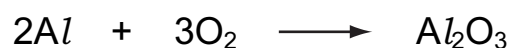
- (i) The electron configuration of an aluminium **atom** is **2,8,3**.

Explain why the electrical charge of an aluminium **ion** is +3.

.....

 [2]

- (ii) A student suggested the symbolic equation below for the formation of aluminium oxide.



State and explain whether or not this equation is balanced.

.....
.....
..... [2]

- (c) The firework mixture contained in the firework in Fig. 6.1 contains the compound potassium perchlorate, $KClO_4$.

When potassium perchlorate is heated, a colourless gas is given off which re-lights a glowing splint.

Suggest why the firework mixture needs to contain potassium perchlorate.

.....
.....
..... [2]

7 (a) State which type of electromagnetic wave

(i) can be detected by the human eye, [1]

(ii) is used in a remote control for a television, [1]

(iii) is strongly absorbed by the water in cells. [1]

(b) Three types of nuclear radiation are alpha, beta and gamma. Each of these can be identified by its behaviour in electric and magnetic fields.

Describe how you could identify alpha, beta and gamma radiations by their deflections in an electric field.

Explain your answer. You may use a diagram to help your explanation.

.....

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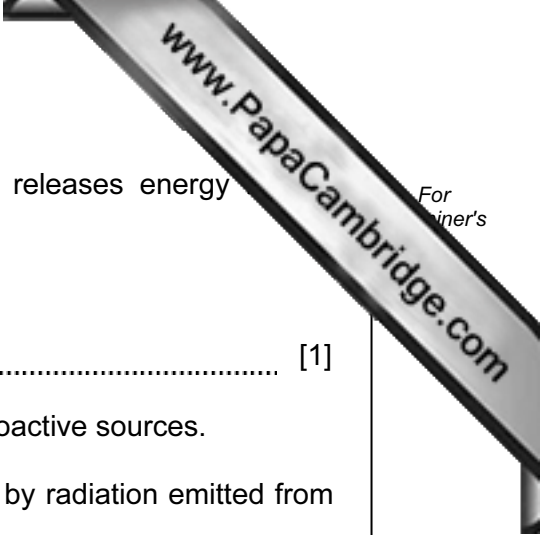
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.....

.....

..... [5]



(c) In a nuclear power station, nuclear fuel such as uranium releases energy through the process of nuclear fission.

(i) State what happens to the uranium atoms.

..... [1]

(ii) At a nuclear power station, technicians work close to radioactive sources.

State **one** way in which these workers could be harmed by radiation emitted from radioactive sources.

.....
..... [1]

(iii) State **two** ways in which these workers could be protected from the radiation.

1
2 [2]

8 Fig. 8.1 shows the male reproductive system.

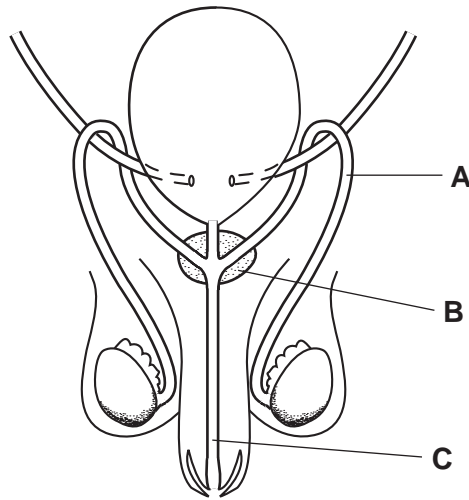


Fig. 8.1

(a) (i) State the functions of parts **A**, **B** and **C**.

- A**
- B**
- C** [3]

(ii) On Fig. 8.1, use a label line and the letter **S** to indicate where male gametes are made. [1]

(b) Describe **three** ways in which human male gametes differ from human female gametes.

- 1
- 2
- 3 [3]

(c) Male gametes and female gametes have a haploid nucleus.

Explain why it is important that gametes have a haploid nucleus.

-
-
-
- [2]

(d) HIV is the virus that causes AIDS. HIV can be passed from one person to another during sexual intercourse. For
Trainer's

Outline how HIV affects the immune system of a person with HIV/AIDS.

.....

.....

.....

..... [2]



- 9 In 1774 the chemist Carl Scheele reacted concentrated hydrochloric acid with manganese dioxide. One of the products of this reaction was a pale green gas which Scheele believed to be a compound containing oxygen.

All attempts by Scheele and other chemists to decompose this green gas were unsuccessful. In 1810 the green gas was named chlorine.

- (a) Explain which information in the passage above suggests that chlorine is an element.

.....

 [2]

- (b) Chlorine is produced in the chemical industry by electrolysis.

A simplified diagram of one type of electrolysis cell used to produce chlorine is shown in Fig. 9.1.

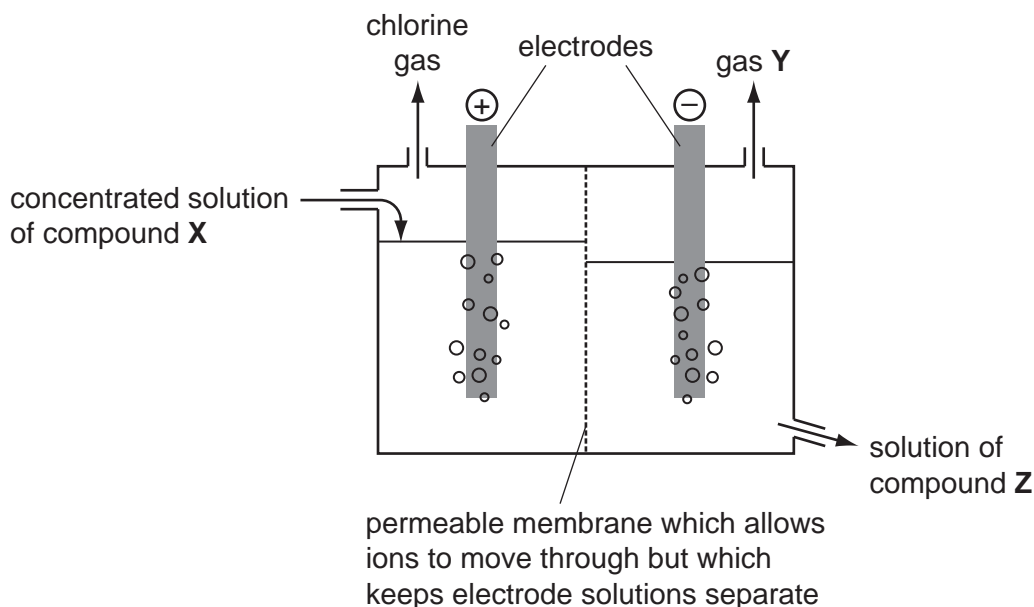


Fig. 9.1

- (i) Name substances X, Y and Z in Fig. 9.1.

X

Y

Z

[3]

- (ii) Fig. 9.2 shows how the electrons are arranged in a chlorine atom.

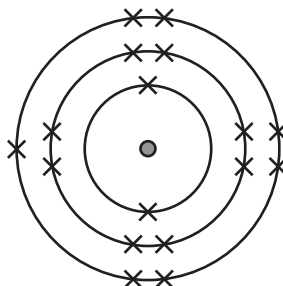


Fig. 9.2

In chlorine gas, the atoms form molecules which have the formula, Cl_2 .

Draw a diagram to show how the **outer** electrons are arranged in a molecule of chlorine.

[2]

- (c) A student plans to produce some chlorine gas by repeating the reaction used by Scheele. She researches the balanced symbolic equation for the reaction and finds that it is



The student decides to react 1.74 g of manganese dioxide with excess hydrochloric acid.

- (i) Calculate the number of moles of manganese dioxide in 1.74 g.

Show your working.

..... [2]

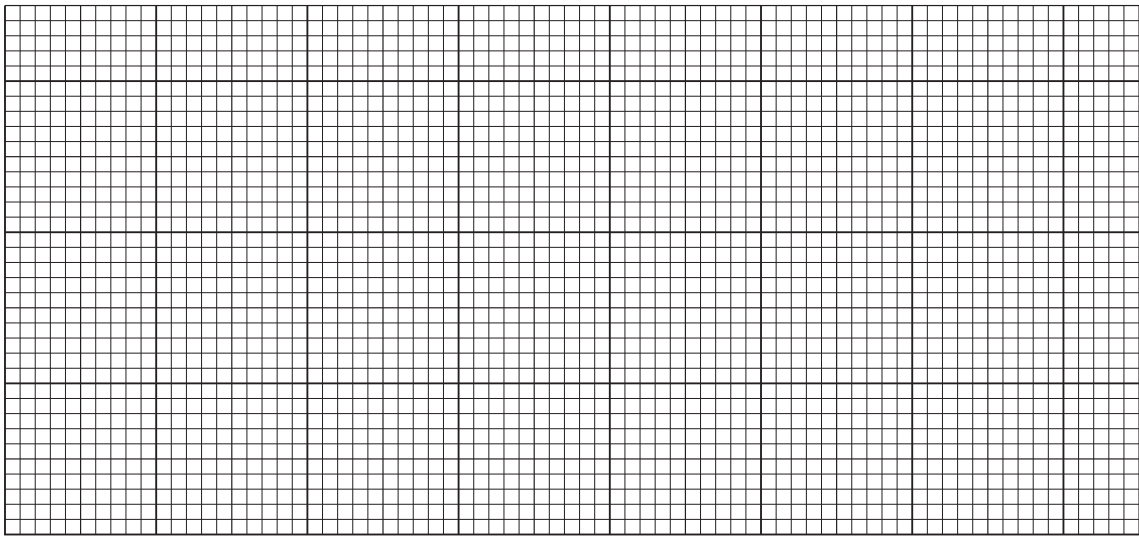
- (ii) Calculate the volume of chlorine gas, measured at room temperature and pressure, which the student might expect to be produced in her experiment.

The volume of one mole of chlorine, measured at room temperature and pressure, is 24 dm^3 .

Show your working.

..... [3]

- 10 (a) On the grid below, draw a wave with an amplitude of 2 cm and a wavelength of 4 cm.
On your diagram, clearly label the amplitude and the wavelength.



[3]

- (b) (i) Two sound waves, **A** and **B**, have the same frequency. **A** has a greater amplitude than **B**.

What difference would you hear?

..... [1]

- (ii) Two sound waves, **X** and **Y**, have the same amplitude but **X** has a greater frequency than **Y**.

What difference would you hear?

..... [1]

- (iii) The speed of sound was calculated for sound passing through a solid, a liquid, a gas and a vacuum.

The values recorded were

0 m/s 330 m/s
1500 m/s 5000 m/s.

Write the values in the correct boxes in Table 10.1.

Table 10.1

	<u>speed of sound</u> m/s
vacuum	
solid	
liquid	
gas	

[2]

- (iv) Sound travels through the air by a series of compressions and rarefactions.

Explain what is meant by *compressions* and *rarefactions*. You may use a diagram to help your explanation.

.....

.....

..... [2]

(c) Energy travels to the Earth from the Sun.

State whether this transfer of energy is by conduction, convection or radiation.

Explain your answer.

.....
.....
..... [2]

(d) Many bush fires are caused by pieces of glass that have been carelessly thrown away.

Fig. 10.1 shows parallel rays of light passing through a piece of glass. The piece of glass acts as a lens and focuses the light on the ground.

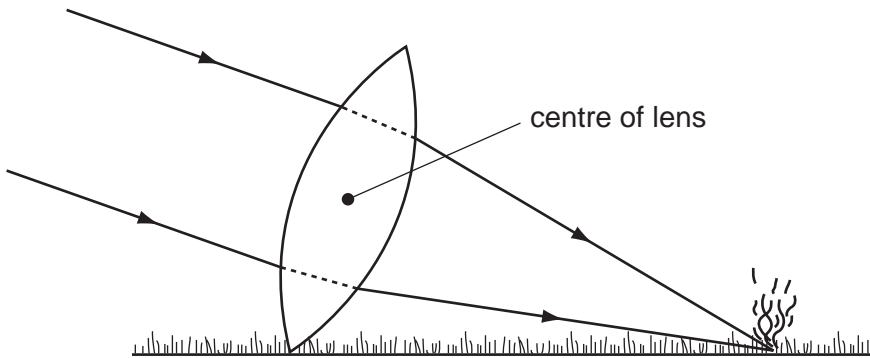


Fig. 10.1

(i) On Fig. 10.1, use the letter **P** to label the principal focus of the piece of glass. [1]

(ii) Measure the focal length of the piece of glass in Fig. 10.1.

..... mm [1]

(iii) The glass acting as a lens produces a real image of the Sun.

Explain what is meant by the term *real image*.

.....
..... [1]

11 Humans require a wide range of nutrients to provide a balanced diet.

(a) List **two** groups of **organic** substances that humans require in their diet.

1

2

[2]

(b) Outline the symptoms that a person may develop if their diet is deficient in

(i) vitamin D,

..... [1]

(ii) iron.

..... [1]

(c) Describe the use of microorganisms in the manufacture of yoghurt.

.....

.....

.....

.....

.....

..... [3]

- 12 (a) (i) Name the **two** elements which are combined together in most of the compounds found in petroleum (crude oil).

1

2

[1]

- (ii) Draw **four** straight lines to connect each process or reaction in the left hand column with its meaning in the right hand column.

type of process or reaction	process or reaction
reaction that produces ethane from ethene and hydrogen	addition
reaction that causes protein molecules to break up into amino acids	catalytic cracking
reaction that produces unsaturated compounds	fractional distillation
process that simplifies a complex mixture	hydrolysis

[2]

- (b) Fig. 12.1 shows apparatus that a student uses to investigate what happens when gaseous decane, $C_{10}H_{22}$, is heated in the presence of a catalyst.

The catalyst is made of small pieces of aluminium oxide which are heated strongly.

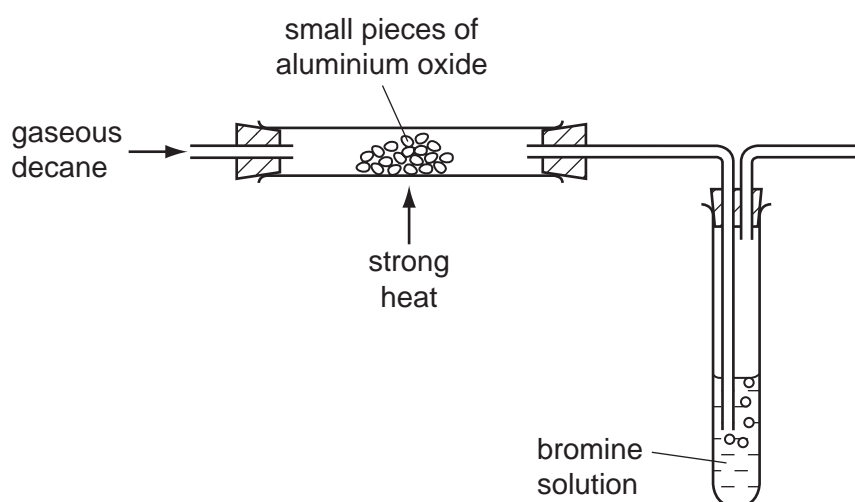


Fig. 12.1

When the gaseous decane passes through the heated catalyst, the solution of bromine rapidly changes colour from orange to colourless.

- (i) Explain why this observation shows that decane has undergone a cracking reaction.

.....

 [3]

- (ii) Explain why the products of the reaction do not include any aluminium compounds.

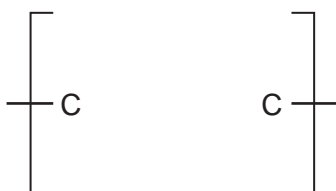
.....
 [1]

- (iii) Suggest why the catalyst needs to be heated.

.....
 [1]

- (c) When ethene, C_2H_4 , is heated and pressurised in the presence of a catalyst, it is converted into a white compound which becomes solid when it cools.

- (i) Complete the diagram below to show a small section of one of the molecules in the white solid.



[2]

- (ii) Suggest why it is **not** possible to state an exact value of the relative molecular mass of the molecules in the white solid.

.....
 [1]

DATA SHEET
The Periodic Table of the Elements

		Group																			
		I	II	III	IV	V	VI	VII	VIII	IX	X										
		1 H Hydrogen 1																			
7	9	Li Lithium 3	Be Beryllium 4									He Helium 2									
23	24	Na Sodium 11	Mg Magnesium 12									Ne Neon 10									
39	40	K Potassium 19	Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36		
85	88	Rb Rubidium 37	Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	101 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	115 Sn Tin 50	119 Sb Antimony 51	122 Te Tellurium 52	128 I Iodine 53	127 Xe Xenon 54	131 Xe Xenon 54		
133	137	Cs Caesium 55	Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86	210 Rn Radon 86		
87	226	Fr Francium 87	Ra Radium 88	227 Ac Actinium 89																	
												*58-71 Lanthanoid series †90-103 Actinoid series									
		a		X		b		a = relative atomic mass X = atomic symbol b = proton (atomic) number													
		140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	144 Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	175 Lu Lutetium 71	175 Lu Lutetium 71	175 Lu Lutetium 71	175 Lu Lutetium 71	175 Lu Lutetium 71	175 Lu Lutetium 71
		232 Th Thorium 90	232 Pa Protactinium 91	238 U Uranium 92	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103	238 Lr Lawrencium 103	238 Lr Lawrencium 103	238 Lr Lawrencium 103	238 Lr Lawrencium 103	238 Lr Lawrencium 103	238 Lr Lawrencium 103

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).